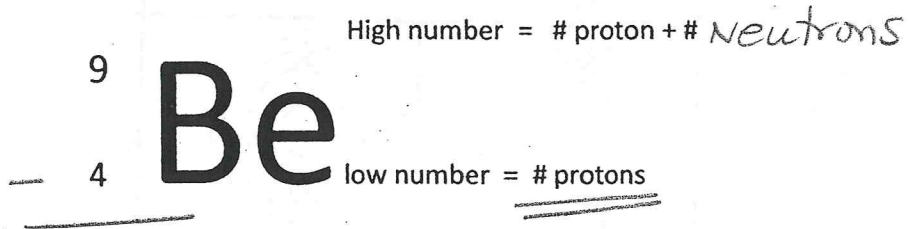
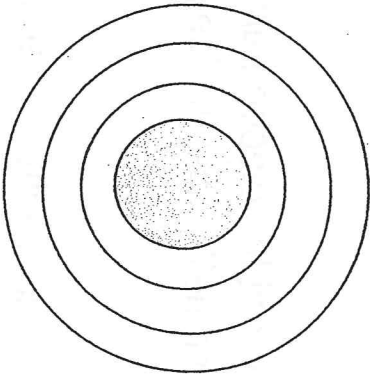


Bohr-Rutherford Atom – How to draw these diagrams

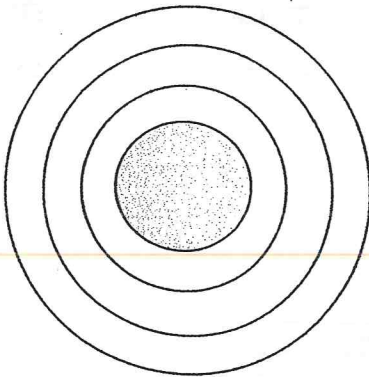
Atomic number = # protons in the \_\_\_\_\_. This is the \_\_\_\_\_ number and is also its position on the periodic table. ~~It is also called the~~

Mass number = # protons + # neutrons. This is the \_\_\_\_\_ number. To find how many neutrons you find the difference (subtract)



(5) So... high number – low number = # neutrons

# electrons = \_\_\_\_\_



O

Sub-atomic Particle	Mass (atomic units)	Charge
Proton		
Neutron		
Electron		

# Electron Arrangement



Analysis:

- Columns:** Look at the electron arrangements in a vertical column. What is the same? What is different?
- Rows:** Look at the electron arrangements in a horizontal row. What is the same? What is different?